Element in period 4 and group 7A.-->Br

Element on the periodic table with the largest atomic radius. 🡪 Fr

The most reactive nonmetal. 🡪 F

The most reactive metal.--> Fr

Element on the periodic table with the smallest atomic radius.--> He

The element having 16 protons.-> S

The element with 3 electrons in the outermost 3rd energy level.--> Al

The only nonmetal on the left side of the table.--> H

The metalloid having 2 energy levels.--> B

A noble gas that does NOT have 8 electrons in the outer level.--> He

Any element in group 1B.--> Cu, Ag, or Au

The halogen in period 5.--> I

The alkali metal having 4 energy levels.--> K

The element that has 6 energy levels and forms a 2+ ion.--> Ba

The only metal that is a liquid at room temperature.--> Hg

The only nonmetal that is a liquid at room temperature.-> Br

Any element in the actinide series.--> Bottom row below the main table

Any element in the lanthanide series.-->Top row below the main table

Element in period 3 with the smallest radius.--> Ar

Element in period 6 with the largest radius.-->Cs

Element in group 5A having the highest ionization energy.-->N

Element in group 3A with the lowest ionization energy.--> Uut

Element having 26 electrons.--> Fe

Element in period 3 that gains 1 electron in order for form a filled outer level.--> Cl

The nonreactive element that has the largest radius.--> Uuo

The noble gas that is in the same period as the lanthanide series.-->Rn

Nonmetal in group 4A.-->C

Alkaline earth metal having the lowest ionization energy.--> Ra

Element in group 6A that is a gas.--> O

Element that is named after the scientist who developed the modern atomic table.--> Md

Element whose configuration is [Ne]3s23p5.--> Cl

Metal in group 4A with the smallest radius.-->Sn

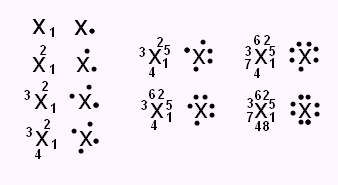
Nonmetal in group 5A with the highest ionization energy.-->N

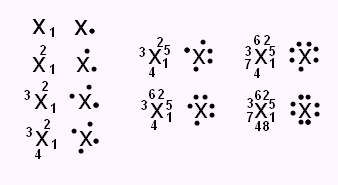
Metalloid that has 5 valence electrons.-->P or As

Largest element that forms a 2- ion.--> Lv

Smallest metal that loses 1 electron to form a filled outer level.--> Li

Element that has a half-filled 4d sublevel.--> Tc

Element that has the following dot notation: 🡪 any element in column 16 (6A)

Element with the following dot notation: 🡪 any element in column 15 (5A)

Element in the “s” block.--> any element in group 1A or 2A (column 1 or 2)

Element in the “p” block.--> any element in groups 3A-8A (column 13-18)

Element in the “d” block.--> any element in columns 3-12

Element in the “f” block.--> any element in the two rows below the main part of the table

Element that is a representative element.--> any elements in columns 1, 2, or 13-18 (“A” groups)

Element that is a transition element.--> any element in the “d” block (columns 3-12)

Nonreactive element that has 5 energy levels.--> Xe

Element in period 3 group 5A.-->P

Element whose configuration ends 5p5.-->I

Element in period 6 that forms an ion with a 3+ charge.-->Tl

Element that gains 2 electrons to become stable and has an atomic number less than 12.--> O

Element that has one more proton than tungsten.--> Re

Element that has one more electron than the bromine ion.--> Rb

Element whose electron configuration ends 3p4.--> S

Element that has similar chemical properties as barium.--> any element in column 2 (2A)

Element that is isoelectronic with Cl1-.--> Ar

Element that forms a 1+ ion that is isoelectronic with argon.--> K

Element with one less electron that silver.--> Pd

Element that is named after a planet.--> Np, Pu, Hg, U

Element in period 5 group 5B.--> Nb