Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date \_\_\_\_\_\_ Hour \_\_

## *Mass, Volume, and Density*

|  |  |
| --- | --- |
| 1. Study the matter shown in Figure 1. Each dot represents a particle of matter. [*Assume the particles are uniformly distributed throughout each object, and particles of the same size have the same mass*.]   * 1. In the table below, show how the masses, volumes, and densities of A and B compare by adding the symbol **<, >, or =** to the statement in the second column.   2. Explain your reasoning for each answer in the last column. |  |

|  |  |  |
| --- | --- | --- |
| **Property** | **Relationship** | **Reasoning** |
| Mass | A \_\_<\_\_ B | Fewer particles present in A |
| Volume | A \_\_<\_\_ B | B is larger in size |
| Density | A \_\_<\_\_ B | B has a greater mass to volume ratio |

**FIGURE 2**

**A B**

**C**

2. Study the matter in Figure 2. [*Assume the particles are uniformly distributed throughout each object, and particles of the same size have the same mass*.]

a. In the table below show how the masses, volumes, and densities compare by adding the symbol **<, >, or =** to the statement in the second column.

b. Explain your reasoning for each answer in the last column.

|  |  |  |
| --- | --- | --- |
| **Property** | **Relationship** | **Reasoning** |
| Mass | A \_\_<\_\_ B  A \_\_=\_\_ C | A has fewer particles |
| Volume | A \_\_<\_\_ B  A \_\_=\_\_ C |  |
| Density | A \_<\_\_\_ B  A \_\_<\_\_ C | A has smaller mass to volume ratio |

1. Is object E or object F denser? [*Assume the particles are uniformly distributed throughout each object, and particles with a larger size have a larger mass*.] Explain your reasoning.

**FIGURE 3**

**E F**

**Box on left has greater density. More mass in same volume**

4. In figure 4 below, a graph shows the relationship between mass and volume for two substances, A and B. Use the graph to answer questions about these substances.

**Two Pan Balance**

**A B**



1. You built a simple two-pan balance, shown above, to compare the masses of substances **A** and **B** shown in the graph. What would happen to the balance if you put **equal masses** of **A** and **B** in the two pans? Explain your reasoning.

The balance would remain as is. Balances measure mass and if the samples are equal in mass the balance will remain as is.

1. What would happen to the balance if you put **equal volumes** of **A** and **B** in the two pans? Explain your reasoning. Pan A would drop lower than B. A has a greater density.
2. If you put **10.0 mL of A** in one balance pan, how much **mass of B** would you need in the other pan to make it balance? Explain your reasoning. Approximately 15 g.
3. If you put **35.0 mL of B** in one balance pan, what **volume of A** would you need in the other pan to make it balance? Explain your reasoning. Approximately 12 mL
4. Find the slope of the line for both A and B using correct units. State the physical meaning of the slope for each substance. The slope of each line is the density
5. Water has a density of 1.00 g/mL. **Sketch the line** representing water on the graph in Figure 4.
6. Determine whether substance A and B will sink or float when placed in a bucket of water.

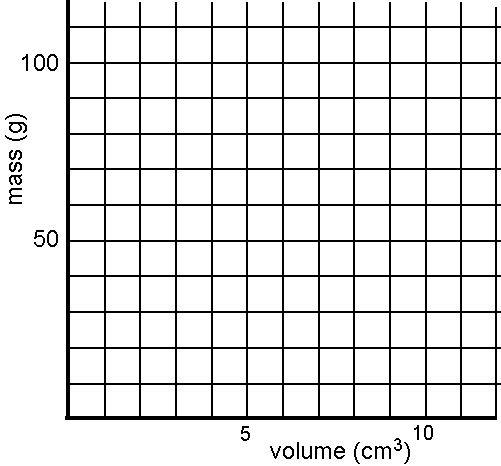
**A**: sink float **B**: sink float (circle correct response)

Defend your answer using the mass-volume graph (Figure 4), and your outstanding understanding of density.

Objects sink when the density is greater than the density of the substance they are placed in.

Refer to the table of densities at the right to answer the following questions.

|  |  |
| --- | --- |
| **Substance** | **Density (g/mL)** |
| Aluminum | 2.70 |
| Titanium | 4.54 |
| Zinc | 7.13 |
| Tin | 7.31 |
| Iron | 7.87 |
| Nickel | 8.90 |
| Copper | 8.96 |
| Silver | 10.50 |
| Lead | 11.35 |
| Mercury | 13.55 |
| Gold | 19.30 |



5. Sketch a graph of mass vs. volume for titanium, copper and mercury.

6. You made some cubes out of each metal in the table. Each measures 2.00 cm on every side (all except mercury). Why can’t you make a cube of mercury? Mercury is a liquid at room temp

a. What is the volume of each cube in **cm3**? in **mL**? (Show your calculations)

V = \_\_8.00\_\_\_\_ cm3 V = \_\_8.00\_\_\_\_ mL

b. Find the mass of these metal cubes: (Show your work below)

lead cube \_\_90.8 g\_\_\_\_\_\_\_\_\_\_\_\_

nickel cube \_\_71.2 g\_\_\_\_\_\_\_\_\_\_\_\_

zinc cube \_\_\_57.0 g\_\_\_\_\_\_\_\_\_\_\_

7. Alicia’s cheapskate boyfriend gave her a ring he claims is 24 carat gold. Alicia is skeptical. After chemistry class, she measures the mass of the ring, finds the volume of the ring by water displacement, and then calculates the density of the ring. Should she treasure the ring as his first truly generous gift to her, or throw it at him the next time he walks by? **Show your calculations below and defend your answer using the density you calculated**.

DATA:

Mass: 15.28 g

Final volume: 43.7 mL

Initial volume: 42.2 mL

Volume of ring: \_\_\_\_1.5 mL\_\_\_\_\_\_

Density: \_\_\_10.2\_\_\_\_\_\_\_ Not gold

9. A student filled a graduated cylinder with water and read the meniscus at 25.8 mL. The student then dropped a solid material into the graduated cylinder and the water level rose to 35.9 mL. If the solid material had a density of 2.99 g/mL, determine the mass of the solid object.

**35.9-25.8 = 10.1 mL x 2.99 = 30.2 g**

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*Density**Homework*

Refer to the densities below to answer the following questions.

Iron 7.87 g

1 mL

Copper 8.96 g

1 mL

Titanium 4.54 g

1 mL

You have some iron wire, copper wire, and titanium wire (all the same gauge, or diameter). Your lab group measured out a length of wire that is exactly 10.00 g for each type of metal wire.

a) Which of these 3 metal wires would be the longest? ***Explain your reasoning***.

b) Which of these 3 metal wires would be the shortest? ***Explain your reasoning.***

*Extra Credit*

d) If every 1**.**0 cm length of the titanium wire has a mass of 0.15 g, how long would the 10.00 g wire be? (Hint: write a **conversion ratio** for the two quantities you are working with)

**Answers**: #4 to end (quantitative problems, primarily)

4a) Balance would tip to the right: mA>mB for the same volume

b) DA=60g/45mL = **1.3g/mL**; DB = 32g/60mL = **0.53g/mL**

Each mL of substance has a mass of (1.3 or 0.53) grams.

c) mA = mB when it is balanced. mA = 1.3 g/mL \* 10.0 mL = 13 g A => **mB = 13 g**

d) 35 mL B \* 0.53g/mL = 18.6 g B = 18.6 g A

18.6g A \*1mL/1.3 g = **14.3 mL A**

e) (sketch line on graph w/slope of 1.0g/1mL)

f) **A will sink**, (dA>dwater); **B will float**, (dB<dwater)

5. Copper; Mercury; Titanium

6 a) volume = 2.0 cm \* 2.0 cm \* 2.0 cm = 2.03 cm3 = **8.0 cm3 = 8.0 mL**

b) **Volume \* Density = mass**

8.0 mL \* 11.35 g/mL = 90.8 g Pb

\* 8.90 g/mL = 70.2 g Ni

\* 7.13 g/mL = 57.0 g Zn

7) Density of ring =  = 10.2 g/mL; Looks like Pb/Cu mix!! Throw it at him!!

**Extra Credit:**

a, c) Longest: Titanium (needs more volume (=length) than the others for 10g mss because its density is smallest)

b, c) Shortest: Copper (needs less volume (=length) than the others for 10 g mass because its density is the greatest)

d) 



v = r2l



